

The Imitation Of Gestures A Technique For Studying The Body Schema And Praxis Of Children Three To Six Years Of Age Ir Egrave Ne L Eacute Zine

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this book published with two online only appendices is designed to show and discuss another facet of Stosch that would argue with the dense mythology of a spy hoarder and libertine clouding the true nature of his accomplishments as an antiquarian collector patron and scholar this is possible due to discovery and study of a substantial part of Stosch's previously considered lost enormous paper museum of gems the artists including Pier Leone Ghezzi Girolamo Odam Bernard Picart Antonio Maria Zanetti Markus Tuscher Theodor Netscher Georg Martin Preißler Johann Justin Preißler and Johann Adam Schweickart tirelessly worked in a studio organised by Stosch on the faithful documentation of vast numbers of engraved gems made for a variety of purposes they expose Stosch's crucial role in the creation and transfer of knowledge that contributed immensely to the transformation of eighteenth century antiquarianism towards a more scholarly archaeological science

an engraved landscape is a contextual analysis of a substantial new corpus of engravings from the wadi al ajal situated in the central saharan region of south west libya the wadi is renowned as the heartland of the garamantian civilization which emerged from local mobile pastoral communities in the 1st millennium bc and dominated trans saharan trade and politics for over a thousand years extensive archaeological and palaeoenvironmental investigations in recent years have provided detailed insight into the later prehistory and protohistory of the wadi and surrounding areas however prior to the fieldwork detailed in this work only a handful of carvings had been recorded in the wadi this work is based on systematic survey conducted between 2004 and 2009 which recorded around 2 500 previously unknown or unpublished engraved and inscribed rock surfaces all forms of engraving whether figurative or surface markings were viewed as significant residues of human interaction with the rock surface and were recorded the resulting database provides an opportunity to analyze the engravings in relation to their changing physical and cultural contexts and the discussion offers a fresh interpretation of saharan rock art based on this substantial new evidence an engraved landscape also captures in detail a unique heritage resource that is currently inaccessible and threatened this record of the fragile engravings provides an important source of information for researchers and students

this book studies small but highly captivating artworks from antiquity engraved gemstones these objects had multiple applications and the images upon them captured snapshots of people s beliefs ideologies and everyday occupations they provide a unique perspective on the propaganda of roman political leaders especially octavian augustus

at head of title comune di udine assessorato alla cultura civici musei di storia ed arte gabinetto numismatico

first published in 1998 engraved on steel focuses on engraving and engravers exploring the use of steel engraving in both the decorative arts and in printing basil hunnisett also describes the context of the steel engraver s work the processes by which steel engraving became one of the most widely used forms of printing in the 19th century are described in detail as the developments in the print industry paper manufacture and publishing that determined its history the activities of print publishers are also examined including those of art unions

a comprehensive study of the depictions of animals and their significance on greek and roman gems the work examines the associations between animal depictions and the type of gemstone and its believed qualities the study also compares the representation of animals on gems to other larger media and analyses the differences

reprint of the original first published in 1876

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Understanding the KCL Structure: A Simple Guide

Potassium chloride (KCl), a common salt, is more than just a laboratory reagent; it's a fundamental example of ionic bonding and crystal structures. Understanding its structure helps us grasp the principles behind many other ionic compounds and their properties. This article simplifies the concept of KCl's structure, making it accessible even without a strong background in chemistry.

1. Ionic Bonding: The Foundation of KCL Structure

KCl's structure arises from the fundamental interaction between potassium (K) and chlorine (Cl) atoms. Potassium, an alkali metal, readily loses one electron to achieve a stable electron configuration, becoming a positively charged ion (K^+). Chlorine, a halogen, readily gains one electron to achieve a stable configuration, becoming a negatively charged ion (Cl^-). This electron transfer constitutes an ionic bond – a strong electrostatic attraction between oppositely charged ions. The powerful attraction between K^+ and Cl^- is the driving force behind the formation of the KCl crystal lattice.

2. The Cubic Crystal Lattice: Arrangement of Ions

The K^+ and Cl^- ions don't arrange themselves randomly. Instead, they form a highly ordered three-dimensional structure called a cubic crystal lattice. Imagine a cube; within this cube, K^+ and Cl^- ions alternate systematically. A K^+ ion is surrounded by six Cl^- ions, and conversely, a Cl^- ion is surrounded by six K^+ ions. This arrangement maximizes the electrostatic attraction between the positive and negative ions while minimizing repulsion between ions of the same charge. This specific arrangement is known as a face-centered cubic (FCC) lattice, where both K^+ and Cl^- ions occupy the lattice points.

3. Coordination Number and Unit Cell: Defining the Structure

The coordination number describes the number of ions of opposite charge surrounding a central ion. In KCl, both K^+ and Cl^- have a coordination number of 6. This signifies the strong interaction between neighboring ions. The unit cell is the smallest repeating unit of the crystal lattice. For KCl, the unit cell is a cube containing one K^+ ion and one Cl^- ion at the corners and face centers. All the properties of the crystal are defined by this repeating unit cell.

4. Properties arising from the KCL Structure:

The ordered arrangement of ions in the KCl lattice directly influences its properties. For example: High melting and boiling points: The strong electrostatic forces between K^+ and Cl^- require significant energy to overcome, resulting in high melting and boiling points. This is typical of ionic compounds. Solubility in water: Water, being a polar solvent, can effectively interact with the charged ions, overcoming the electrostatic forces and dissolving the KCl crystal. Electrical conductivity: Solid KCl is a poor conductor of electricity because the ions are fixed in the lattice. However, molten KCl or a KCl solution conducts electricity well because the ions are free to move. Brittleness: Applying stress to the crystal can cause layers of ions to shift, leading to repulsion between ions of the same charge. This repulsion overcomes the attraction, resulting in the crystal fracturing – hence the brittleness.

5. Real-World Applications: From Medicine to Industry

KCl's unique structure and properties lead to diverse applications: Medicine: KCl is crucial for maintaining electrolyte balance in the body. It's used to treat potassium deficiency (hypokalemia). Food industry: KCl serves as a salt substitute in low-sodium diets. Fertilizers: Potassium is an essential plant nutrient, and KCl is a major component of many fertilizers. Industry: KCl is used in various industrial processes, including the production of potassium hydroxide (KOH) and other potassium compounds.

Key Insights:

KCl's structure is a direct result of ionic bonding and the electrostatic attraction between K^+ and Cl^- ions. The crystal lattice is a highly ordered arrangement maximizing attraction and minimizing repulsion. The properties of KCl are directly related to its crystal structure and ionic bonding.

FAQs:

1. What is the difference between a crystal lattice and a unit cell? A crystal lattice is the overall repeating 3D arrangement of ions. A unit cell is the smallest repeating unit of that lattice. 2. Why is KCl brittle? The repulsion between similarly charged ions when layers shift under stress causes the crystal to fracture. 3. Is KCl soluble in oil? No, KCl is not soluble in oil because oil is a nonpolar solvent, and KCl is an ionic compound. 4. How does KCl conduct electricity? Only when molten or dissolved in water, where ions are free to move and carry charge. 5. What is the role of KCl in fertilizers? Potassium, derived from KCl, is an essential macronutrient for plant growth, promoting healthy root development and overall plant health.

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